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Holt Chemistry Chapter 6 Covalent Compounds [EPUB] Holt Chemistry Chapter 6 Covalent Compounds. covalent bond. molecular orbital. what is created when two atomic orbital.... bond length. a bond formed when atoms share one or more pairs of electrons. the region of high probability that is occupied by an individu.... a molecular orbital.

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Holt Chemistry Chapter 6. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. MrsCrissyDaniel TEACHER. Section 1. Section 2. Section 3. Terms in this set (12) covalent bond. ... A covalent bond in which a shared pair of electrons is held more closely by one of the atoms.(p.194)

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Holt Chemistry Chapter 6. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. eman6044. The questions that she said will be on the test. Terms in this set (22) ... How does a covalent bond differ from an ionic bond? Ionic- Electrons are transferred Covalent- Electrons are shared.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

6 Chemical Bonding

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Holt Chemistry - Chapter 6 Vocabulary. the region of high probability that is occupied by an individual electron as it travels with a wavelike motion in the three dimensional space around one of two or more associated nuclei. This activity was created by a Quia Web subscriber.

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Holt Chemistry Chapter 6. covalent bond. molecular orbital. bond length. bond energy. A bond formed when atoms share one or more pairs of electrons..... The region of high probability that is occupied by an individu.... The distance between two bonded atoms at their minimum potenti....

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6) purely ionic. b. partially ionic. c. polar-covalent. d. nonpolar-covalent. 10. You can estimate the degree to which a bond between two atoms is ionic or covalent by calculating the a. distance between the atoms' nuclei. b difference in the atoms' electronegativities. c. atoms' atomic radii. d. number of atoms in the compound.